**Honors Chemistry Unit Four: Reactions and Moles Assignments**

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| **Week** | **Date** | **Assignments** | **Due** | **Textbook Section** | **Objectives** | **Turn in or stamp?** |
| Q2 Week 5 | 11/17 | HW: Writing Chemical Equations Notes | 11/18 | 10.1 | 1 – 2 | S |
| 11/18  CP | CW: Writing Chemical Equation Practice | 11/18 | 10.1 | 1 - 2 | T |
| HW: Balancing Chemical Equation Notes | 11/19 | 10.1 | 1 – 3 | S |
| 11/19 | CW: Balancing Chemical Equations WS | 11/19 | 10.1 | 1 – 3 | T |
| HW: Types of Reactions Notes and Practice Classifying Reactions Notes | 11/20 | 10.1 | 4 | S |
| 11/20 | CW: Types of Reactions Lab | 11/20 | 10.2 | 1 – 4 | T |
| HW: Single Replacement Reaction Notes | 11/21 | 10.2 | 5 | S |
| Week 6 | 11/23 | CW: Single Replacement Reaction Lab | 11/23 | 10.2 | 1 – 5 | T |
| HW: Solubility Rules Notes and Precipitation Reaction Notes | 11/24 | 10.3 | 6 – 8 | S |
| 11/24 | CW: Double Replacement WS | 11/24 | 10.3 | 6 – 8 | T |
| HW: Net Ionic Equation Notes | 11/30 | 10.3 | 6 – 9 | S |
| Week 7 | 11/30 | CW: Precipitate Lab 2 | 12/1 | 10.3 | 6 – 9 | T |
| HW: Precipitate Lab Analysis |
| 12/1 | CW: Review WS | 12/2 | 10 | 1 – 9 | T |
| HW: Finish Review WS |
| 12/2  CP | **\*\*\*Quiz- Chapter 10\*\*\*\*** | | | | |
| HW: Mole and Avogadro Notes and Molar Mass Notes | 12/3 | 11.1 – 11.2 | 10 – 13 | S |
| 12/3 | CW: Txtbk pg 346 #71 – 75, 91, 99, & 111 | 12/3 | 11.2 | 10 – 13 | T |
| HW: Mole Calculation 1 Notes | 12/4 | 11.3 | 10 – 14 | S |
| 12/4 | CW: Mole Station Lab | 12/4 | 11.3 | 10 – 14 | T |
| HW: %composition notes and Empirical Formula and Mole | 12/7 | 11.4 | 15 – 16 | S |
| Week 8 | 12/7 | CW: Txtbk pg 346 #83, 86, 122, 124, 136, 144, & 147 | 12/7 | 11.4 | 15 – 16 | T |
| HW: Determining Formulas Notes | 12/8 | 11.4 | 15 – 18 | S |
| 12/8 | CW: Empirical Formula Lab | 12/9 | 11.4 | 15 – 16 | T |
| HW: Determining Formulas Practice Problems notes | 12/9 | 11.4 | 15 – 18 | S |
| 12/9 | CW: Txtbk pg 346 #84, 143, 145, 146, 148 – 150, & 160 | 12/9 | 11.4 | 15 – 18 | T |
| HW: Formula of a Hydrate Notes | 12/10 | 11.5 | 19 – 20 | S |
| 12/10 | CW: Formula of a Hydrate Lab | 12/11 | 11.5 | 19 – 20 | T |
| HW: Objectives | 12/14 | 10/11 | 1 – 20 | S |
| 12/11 | CW: Review |  |  |  |  |
| HW: Objectives | 12/14 | 10/11 | 1 – 20 | S |
| Week 9 | 12/14 | **\*\*\*Chapter 11 Quiz\*\*\*** | | | | |
| **\*\*\*Unit 4 Binder Check\*\*\*** | | | | |
| HW: Study for Final (Units 1 – 4) |  |  |  |  |
| 12/15 | CW: Review for Final |  |  |  |  |
| HW: Study for Final (Units 1 – 4) |  |  |  |  |
| 12/18 | **\*\*\*Semester One Final\*\*\*** | | | | |

Objectives:

1. I know the parts of an equation: reactants, products, state symbols, yield sign.
2. I can write chemical equations from the names of the formulas.
3. I know how to balance equations.
4. I know how to classify reaction types: synthesis, decomposition, single replacement, double replacement, and combustion.
5. I can predict the products of a single replacement reaction and use an activity series to decide if there is no reaction between the two reactants.
6. I can predict the products of a double replacement reaction.
7. I have the solubility rules memorized for Nitrates, Acetates, Alkali metals, Ammonium, Chlorides, Bromides, Iodides, Sulfates, and Hydroxides.
8. I can use solubility rules to identify the solid produced in a double replacement reaction.
9. I can identify the spectator ions in a molecular (total ionic) equation and cancel them to write a net ionic equation.
10. I know the value of Avogadro’s number.
11. I can use Avogadro’s number to convert between number of particles and moles.
12. I know how to determine the molar mass of an element or compound.
13. I can use molar mass to convert between mass and moles.
14. I can convert between number of particles and mass.
15. I can calculate the percent composition of each element in a compound.
16. I know the difference between empirical formula and molecular formula.
17. I can determine the empirical formula of a compound using percent composition or mass.
18. I can determine the molecular formula of a compound if given the molar mass of the compound and information to determine the empirical formula.
19. I can determine the formula for a hydrate from laboratory data.
20. I can explain what a hydrate is and how its name reflects its composition.